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4.2 For the reaction, $2A + B \longrightarrow A_2B$, the rate = $k[A][B]^2$ with $k = 2.0 \times 10^{-6} \text{ mol}^{-2} \text{ L}^2 \text{ s}^{-1}$. Calculate the initial rate of the reaction when $[A] = 0.1 \text{ mol L}^{-1}$, $[B] = 0.2 \text{ mol L}^{-1}$. Calculate the rate of reaction after $[A]$ is reduced to 0.06 mol L^{-1} . Sol. Initial rate of reaction = $k[A][B]^2 = (2.0 \times 10^{-6} \text{ mol}^{-2} \text{ s}^{-1})(0.1 \text{ mol L}^{-1})(0.2 \text{ mol L}^{-1})^2 = 8 \times 10^{-9} \text{ mol L}^{-1} \text{ s}^{-1}$.

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In modern theory, it is now referred as coordination number of central metal atom or ion. 9.2 FeSO_4 solution mixed with $(\text{NH}_4)_2\text{SO}_4$ solution in 1:1 molar ratio gives the test of Fe^{2+} ion but CuSO_4 solution mixed with aqueous ammonia in 1:4 molar ratio does not give the test of Cu^{2+} ion. Explain why. Ans. FeSO_4 solution mixed with $(\text{NH}_4)_2\text{SO}_4$ solution in 1 : 1 molar ratio forms a double ...

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3.7 Define conductivity and molar conductivity for the solution of an electrolyte. Discuss their variation with concentration. Sol. Conductivity of a solution is defined as the conductance of a solution of 1 cm length and having 1 sq. cm as the area of cross-section. Molar conductivity of a solution at a dilution (V) is the conductance of all the ions produced from one mole of the electrolyte ...

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12 Chemistry Solid State NCERT In Text Solution part 1

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